WPHS .	Honors	Chemistry
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WFHS • Fioliors Chemistry	Period 7 Date 12/16/15 Points 4/4
Unit 4 Bonding Quiz A	Period Y Date (L/116/11) Points 4/4
1) Elements from opposite sides of the periodic table	tend to form
A) covalent compounds B) onic compounds C) compounds that are gaseous at room temperature D) homonuclear diatomic compounds	
E) covalent compounds that are gaseous at room tem	perature
2) The ability of an atom in a molecule to attract elec-	otrons is best quantified by the
A) paramagnetism B) diamagnetism C) electronegativity D) electron change-to-mass ratio E) first ionization potential	
3) The type of compound that is most likely to contain	in a covalent bond is
A) one that is composed of a metal from the far left of the periodic table B) a solid metal C) one that is composed of only nonmetals D) held together by the electrostatic forces between of	of the periodic table and a nonmetal from the far right of
E) There is no general rule to predict covalency in bo	
4) When a metal and a nonmetal react, thegain electrons.	tends to lose electrons and the tends to
A) metal, metal B) nonmetal, nonmetal C) netal, nonmetal D) nonmetal, metal	
E) None of the above, these elements share electrons	25
Multiple Choice Answers: (Please use CAPITAL).	L letters)

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Period		Date _	1_	1_	Points	2	_/4

Unit 4 Bonding Quiz
Elements from opposite sides of the periodic table tend to form
A) covalent compounds
B) ionic compounds
Exampounds that are gaseous at room temperature
D'homonuclear diatomic compounds
E)covalent compounds that are gaseous at room temperature
2) The ability of an atom in a molecule to attract electrons is best quantified by the
A) paramagnetism
B) diamagnetism
(C) electronegativity
D) electron change-to-mass ratio
E) first ionization potential
3) The type of compound that is most likely to contain a covalent bond is
A) one that is composed of a metal from the far left of the periodic table and a nonmetal from the far right of the periodic table
B) a solid metal
(2) one that is composed of only nonmetals
D' held together by the electrostatic forces between oppositely charged ions
By There is no general rule to predict covalency in bonds.
4) When a metal and a nonmetal react, the tends to lose electrons and the tends to gain electrons.
A) metal, metal
B) nonmetal, nonmetal
(C))metal, nonmetal
D) nonmetal, metal
E) None of the above, these elements share electrons
Multiple Choice Answers: (Please use CAPITAL letters)
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