

Unit 4 Bonding Quiz *ReTake*

1) Elements from opposite sides of the periodic table tend to form _____.

- A) covalent compounds
 B) ionic compounds
 C) compounds that are gaseous at room temperature
 D) homonuclear diatomic compounds
 E) covalent compounds that are gaseous at room temperature

2) The ability of an atom in a molecule to attract electrons is best quantified by the _____.

- A) paramagnetism
 B) diamagnetism
 C) electronegativity
 D) electron change-to-mass ratio
 E) first ionization potential

3) The type of compound that is most likely to contain a covalent bond is _____.

- A) one that is composed of a metal from the far left of the periodic table and a nonmetal from the far right of the periodic table
 B) a solid metal
 C) one that is composed of only nonmetals
 D) held together by the electrostatic forces between oppositely charged ions
 E) There is no general rule to predict covalency in bonds.

4) When a metal and a nonmetal react, the _____ tends to lose electrons and the _____ tends to gain electrons.

- A) metal, metal
 B) nonmetal, nonmetal
 C) metal, nonmetal
 D) nonmetal, metal
 E) None of the above, these elements share electrons

Multiple Choice Answers: (Please use CAPITAL letters)

1. B
 2. C
 3. C
 4. C

Unit 4 Bonding Quiz

1) Elements from opposite sides of the periodic table tend to form _____.

- A) covalent compounds
 B) ionic compounds
 C) compounds that are gaseous at room temperature
 D) homonuclear diatomic compounds
 E) covalent compounds that are gaseous at room temperature

2) The ability of an atom in a molecule to attract electrons is best quantified by the _____.

- A) paramagnetism
 B) diamagnetism
 C) electronegativity
 D) electron change-to-mass ratio
 E) first ionization potential

3) The type of compound that is most likely to contain a covalent bond is _____.

- A) one that is composed of a metal from the far left of the periodic table and a nonmetal from the far right of the periodic table
 B) a solid metal
 C) one that is composed of only nonmetals
 D) held together by the electrostatic forces between oppositely charged ions
 E) There is no general rule to predict covalency in bonds.

4) When a metal and a nonmetal react, the _____ tends to lose electrons and the _____ tends to gain electrons.

- A) metal, metal
 B) nonmetal, nonmetal
 C) metal, nonmetal
 D) nonmetal, metal
 E) None of the above, these elements share electrons

Multiple Choice Answers: (Please use CAPITAL letters)

1. E
 2. C
 3. A
 4. C