

Horizon Assessment System

ID: 471849

Test: 2016 Unit Five The Mole Retest

Name : _____

Student ID : _____



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of Questions: 10

Question 1 :

What is the mass in grams of one mole of sulfur dioxide (SO₂)?

A: 48.1 g

B: 64.1 g

C: 80.1 g

D: 96.1 g

Question 2 :

What is the molar mass of Al(NO₃)₃?

A: 57 g/mol

B: 103 g/mol

C: 165 g/mol

D: 213 g/mol

Question 3 :

How many moles of chlorine gas are contained in 9.02×10^{23} molecules?

A: 1.5 moles

B: 2.0 moles

C: 6.02 moles

D: 9.03 moles

Question 4 :

What scientist is credited with the idea of the MOLE?

A: Earnest Rutherford

B: Amedeo Avogadro

C: John Dalton

D: Albert Einstein

Question 5 :

What is the percent composition of Sodium in Sodium sulfide?

A: 59%

B: 44%

C: 1.5%

D: 42%

Question 6 :

What is the empirical formula of the compound with the molecular formula C_9H_{12} ?

A: C_6H_{12}

B: $C_{12}H_{24}$

C: C_3H_4

D: $C_{4.5}H_6$

Question 7 :

What is the total number of neon atoms contained in 20.2 grams of neon gas?

A: 1.01×10^{24}

B: 2.02×10^{24}

C: 3.01×10^{23}

D: 6.02×10^{23}

Question 8 :

How many moles of CH_4 are contained in 96.0 grams of CH_4 ?

A: 3.00 moles

B: 6.00 moles

C: 12.0 moles

D: 16.0 moles

Question 9 :

What is the percent composition of carbon in carbonmonoxide?

A: 76.7%

B: 27.3%

C: 89%

D: 42.9%

Question 10 :

A compound has an EMPIRICAL formula of CH_2 .

If the molecular mass of the compound is close to 98.0 g/mole what is the molecular formula?

A: C_8H_2

B: C_5H_{10}

C: C_7H_{14}

D: $C_5H_{12}O$